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CHAPTER 2

ATOMS, MOLECULES, AND IONS

Problem Categories

Biological: 2.79, 2.80.

Conceptual: 2.31, 2.32, 2.33, 2.34, 2.61, 2.65, 2.71, 2.72, 2.81, 2.82, 2.93, 2.103, 2.113.

Descriptive: 2.24, 2.25, 2.26, 2.49, 2.50, 2.66, 2.70, 2.76, 2.84, 2.85, 2.86, 2.87, 2.88, 2.90, 2.91, 2.92, 2.94, 2.98, 2.101, 2.114.

Environmental: 2.122.

Organic: 2.47, 2.48, 2.69, 2.105, 2.107, 2.108, 2.115.

Difficulty Level

Easy: 2.7, 2.8, 2.13, 2.14, 2.15, 2.16, 2.23, 2.31, 2.32, 2.33, 2.43, 2.44, 2.45, 2.46, 2.47, 2.48, 2.62, 2.91, 2.92, 2.99, 2.100.

Medium: 2.17, 2.18, 2.24, 2.26, 2.34, 2.35, 2.36, 2.49, 2.50, 2.57, 2.58, 2.59, 2.60, 2.61, 2.63, 2.64, 2.65, 2.66, 2.67, 2.68, 2.69, 2.70, 2.71, 2.72, 2.73, 2.74, 2.75, 2.76, 2.79, 2.80, 2.81, 2.82, 2.83, 2.84, 2.85, 2.86, 2.88, 2.89, 2.90, 2.93, 2.94, 2.95, 2.96, 2.98, 2.101, 2.102, 2.103, 2.110, 2.112, 2.114, 2.115.

Difficult: 2.25, 2.77, 2.78, 2.87, 2.97, 2.104, 2.105, 2.106, 2.107, 2.108, 2.109, 2.111, 2.113.

First, convert 1 cm to picometers.

$$? \text{ He atoms} = (1 \times 10^{10} \cancel{\text{ pm}}) \times \frac{1 \text{ He atom}}{1 \times 10^2 \cancel{\text{ pm}}} = 1 \times 10^8 \text{ He atoms}$$

Note that you are given information to set up the unit factor relating meters and miles.

$$r_{\text{atom}} = 10^4 r_{\text{nucleus}} = 10^4 \times 2.0 \cancel{\text{ cm}} \times \frac{1 \cancel{\text{ m}}}{100 \cancel{\text{ cm}}} \times \frac{1 \text{ mi}}{1609 \cancel{\text{ m}}} = 0.12 \text{ mi}$$

For iron, the atomic number Z is 26. Therefore the mass number A is:

$$A = 26 + 28 = 54$$

Strategy: The 239 in Pu-239 is the mass number. The **mass number** (A) is the total number of

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**John Ellis Price, M. David
Haddock, Horace R. Brock**



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