Name _____ Date _____ Class _____

# 10 CHEMICAL QUANTITIES

## Chapter Test A

### A. Matching

*Match each description in Column B to the correct term in Column A. Write the letter of the correct description on the line.*

| | Column A | Column B |
|---|---|---|
| _____ | 1. standard temperature and pressure | a. the volume occupied by a mole of any gas at STP (22.4 L) |
| _____ | 2. molar volume | b. the mass in amus of an atom of an element |
| _____ | 3. representative particle | c. the number of representative particles in a mole of a substance |
| _____ | 4. molar mass | d. the percent by mass of each element in a compound |
| _____ | 5. atomic mass | e. the mass of a mole of a substance |
| _____ | 6. Avogadro's number | f. the formula that gives the lowest whole-number ratio of the elements in a compound |
| _____ | 7. mole (mol) | g. an atom, a formula unit, or a molecule |
| _____ | 8. empirical formula | h. the mass in grams of a mole of any pure substance |
| _____ | 9. percent composition | i. $6.02 \times 10^{23}$ representative particles of a substance |
| _____ | 10. molar mass | j. 0°C and 101.3 kPa |

### B. Multiple Choice

*Choose the best answer and write its letter on the line.*

_____ 11. How many hydrogen atoms are in 4 molecules of isopropyl alcohol, $C_3H_8O$?
a. $4 \times (6.02 \times 10^{23})$
b. 4
c. 32
d. $32 \times (6.02 \times 10^{23})$

_____ 12. The mass of a mole of NaCl is its:
a. molar mass.
b. empirical formula.
c. gram atomic mass.
d. atomic mass.

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# Chapter 10 Chemical Quantities Test Answers

Jonathan Crowe,Tony Bradshaw,Paul M. S. Monk

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**Study Guide/Selected Solutions Manual** Julie R. Frentrup,2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

**Chemical Surety** United States. Department of the Army,1995 *Chemistry for the Biosciences* Jonathan Crowe,Tony Bradshaw,Paul M. S. Monk,2006 Leading students through the essential concepts that are central to understanding biological systems this text uses everyday examples and analogies to build their confidence in an often daunting subject By focusing on the key themes that unify the subject it shows how integral chemistry is to the biosciences **Basic Chemistry** Steven S Zumdahl,1996 Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction Programmed Learning and Individually Paced Instruction Carl H. Hendershot,1973 *College Chemistry Companion* James Kendall,1924 *Basic Concepts of Chemistry, Study Guide* Leo J. Malone,1989-02-21 This Third Edition revised to provide smoother transitions between topics employs a concise yet informal approach to basic chemistry organized to help students employ basic math skills and problem solving strategies Writing style is straightforward and presentation incorporates many concrete analogies to clarify new concepts Includes many illustrative worked examples *Organic Chemistry, Or, The Happy Carbon* I. David Reingold,2007 This Is A Course In Organic Chemistry Yikes Isn T That The Killer Course That Sophomores Around The World Dread Why Are They Teaching It To Us Students Taking Our First Chemistry Course How Will We Survive *Introductory Chemistry* Steven S. Zumdahl,1996 **General, Organic, & Biological Chemistry** Janice Gorzynski Smith,2024-12 My goal in writing this text was to relate the fundamental concepts of general organic and biological chemistry to the world around us and in this way illustrate how chemistry explains many aspects of everyday life A key feature is the use of molecular art to illustrate and explain common phenomena we encounter every day Each topic is broken down into small chunks of information that are more manageable and easily learned Students are given enough detail to understand basic concepts such as how soap cleans away dirt and why trans fats are undesirable in the diet without being overwhelmed *Ebook: Physical Science* Tillery,2016-04-16 Ebook Physical Science **Chemistry** Theodore Lawrence Brown,1988 Manual of Chemistry William Simon,Daniel Base,1923 Chemical Engineering Catalog ,1925 **Study Guide with Selected Solutions** Karen Timberlake,2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems **Pediatric Nursing** Margot R De Sevo,2014-09-12 This volume in the Content

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