

**10****CHEMICAL QUANTITIES****Chapter Test A****A. Matching**

Match each description in Column B to the correct term in Column A. Write the letter of the correct description on the line.

**Column A**

- \_\_\_\_\_ 1. standard temperature and pressure
- \_\_\_\_\_ 2. molar volume
- \_\_\_\_\_ 3. representative particle
- \_\_\_\_\_ 4. molar mass
- \_\_\_\_\_ 5. atomic mass
- \_\_\_\_\_ 6. Avogadro's number
- \_\_\_\_\_ 7. mole (mol)
- \_\_\_\_\_ 8. empirical formula
- \_\_\_\_\_ 9. percent composition
- \_\_\_\_\_ 10. molar mass

**Column B**

- a. the volume occupied by a mole of any gas at STP (22.4 L)
- b. the mass in amu of an atom of an element
- c. the number of representative particles in a mole of a substance
- d. the percent by mass of each element in a compound
- e. the mass of a mole of a substance
- f. the formula that gives the lowest whole-number ratio of the elements in a compound
- g. an atom, a formula unit, or a molecule
- h. the mass in grams of a mole of any pure substance
- i.  $6.02 \times 10^{23}$  representative particles of a substance
- j.  $0^\circ\text{C}$  and 101.3 kPa

**B. Multiple Choice**

Choose the best answer and write its letter on the line.

- \_\_\_\_\_ 11. How many hydrogen atoms are in 4 molecules of isopropyl alcohol,  $\text{C}_3\text{H}_8\text{O}$ ?
  - a.  $4 \times (6.02 \times 10^{23})$
  - b. 4
  - c. 32
  - d.  $32 \times (6.02 \times 10^{23})$
- \_\_\_\_\_ 12. The mass of a mole of NaCl is its:
  - a. molar mass.
  - b. empirical formula.
  - c. gram atomic mass.
  - d. atomic mass.

# Chapter 10 Chemical Quantities Workbook Answers

**Emilie Sanchez**



## **Chapter 10 Chemical Quantities Workbook Answers:**

AQA Key Stage 3 Science Pupil Book 1 Neil Dixon, Carol Davenport, Nick Dixon, Ian Horseywell, 2018-03-19 Ensure that every student develops the maths literacy and working scientifically skills they need to succeed with this skills focused Pupil Book that contains a variety of activities questions and real world examples that are tailored to the Big Ideas and mastery goals of the AQA KS3 Syllabus Develop conceptual understanding with a variety of questions that require students to apply their knowledge to real world scenarios Build working scientifically skills with various Enquiry activities matched to the AQA syllabus Test understanding and measure progress with factual recall questions developed around the ideas of Generalisations Principles and Models Stretch knowledge and understanding with extend tasks linked to higher order thinking skills Compare Evaluate and Predict Bridge the gap between Key Stages 2 and 3 with a focus on maths and enquiry skills and understanding scientific terminology Provides comprehensive support for non specialist or less confident teachers when used in conjunction with the online Teaching Learning resources Written in association with Sheffield Hallam University The Science Education Team within Sheffield Institute of Education SIOE is one of the leading STEM education groups in Europe with a worldwide reputation for knowledge exchange and research SIOE leads national and international STEM education programmes covering curriculum and pedagogical design and development widening participation to traditionally under represented groups and research in science education      **Basic Chemistry** Steven S Zumdahl, 1996

**A Textbook of Physical Chemistry** A. S. Negi, 2007 The Second Revised Edition Of The Book Is Intended To Meet The Requirement Of The Students Of Science Engineering And Other Professional Courses At The Undergraduate Level It Has Been Planned Strictly In Line With The Syllabi Of Various Indian Universities Who Have Adopted The New Ten Plus Two Plus Three Pattern Of Education A New Chapter On Macromolecules Has Been Added Thus Making A Total Of 27 Chapters In The Revised Edition Chapters On Chemical Equilibrium Colligative Properties Atomic Structures Chemical Bonding Have Been Thoroughly Reshuffled And Rewritten Chapter 25 Has Been Rearranged And Divided Into Two Chapters Viz Molecular Spectroscopy And Electrical And Magnetic Properties New Sections Have Been Added To Chapters On Gaseous State Colligative Properties Electrolytic Conduction Ionic Equilibria Chemical Kinetics Atomic Structure And Chemical Bonding Other Chapters Have Also Been Modified And Redesigned The Subject Matter Has Been Given In A Logical Simple And Lucid Language The Main Aim Has Been On Self Learning Some More Diagrams And Illustrations Have Been Added In This Edition For Explaining The Basics And The Fundamentals Of The Subject Conventional Problems In The Earlier Edition Have Been Dropped But General And Objective Type Problems Are Retained A Considerable Number Of Worked Out Problems Have Been Included In Most Of The Chapters These Would Expose The Students To Applications Of Various Concepts And Fundamentals Of The Subject The Revised Text Largely Uses SI Units But CGS Units Have Been Retained In Those Cases Where The SI Units Have Not As Yet Been Fully Appreciated We Have Attempted To Present A Revised Text That Effectively

Provides Clean Accurate And Balanced Views On Various Topics To Grasp The Fundamentals Of The Subject More Clearly Comprehensively And Concretely The Book Should Meet The Requirements Of Students      Quantitative Chemical Analysis, Sixth Edition Daniel C. Harris,2003 For instructors who wish to focus on practical industrial or research chemistry Includes case studies applications boxes and spreadsheet applications      **Chemical Thermodynamics for Process Simulation** Jürgen Gmehling,Michael Kleiber,Bärbel Kolbe,Jürgen Rarey,2019-03-15 The only textbook that applies thermodynamics to real world process engineering problems This must read for advanced students and professionals alike is the first book to demonstrate how chemical thermodynamics work in the real world by applying them to actual engineering examples It also discusses the advantages and disadvantages of the particular models and procedures and explains the most important models that are applied in process industry All the topics are illustrated with examples that are closely related to practical process simulation problems At the end of each chapter additional calculation examples are given to enable readers to extend their comprehension Chemical Thermodynamics for Process Simulation instructs on the behavior of fluids for pure fluids describing the main types of equations of state and their abilities It discusses the various quantities of interest in process simulation their correlation and prediction in detail Chapters look at the important terms for the description of the thermodynamics of mixtures the most important models and routes for phase equilibrium calculation models which are applicable to a wide variety of non electrolyte systems membrane processes polymer thermodynamics enthalpy of reaction chemical equilibria and more Explains thermodynamic fundamentals used in process simulation with solved examples Includes new chapters about modern measurement techniques retrograde condensation and simultaneous description of chemical equilibrium Comprises numerous solved examples which simplify the understanding of the often complex calculation procedures and discusses advantages and disadvantages of models and procedures Includes estimation methods for thermophysical properties and phase equilibria thermodynamics of alternative separation processes Supplemented with MathCAD sheets and DDBST programs for readers to reproduce the examples Chemical Thermodynamics for Process Simulation is an ideal resource for those working in the fields of process development process synthesis or process optimization and an excellent book for students in the engineering sciences      *General Chemistry: Experiment and Theory* Richard H. Eastman,1970      **Transactions of the Faraday Society** ,1968      **Introductory Chemistry** Steven S. Zumdahl,1993      **A Class Book of Physical Chemistry** Thomas Martin Lowry,Samuel Sugden,1929      *The Chemical Evolution of Phosphorus* Enrique Macia-Barber,2019-11-21 Here is a fascinating reader friendly exploration of the phosphorus enigma The volume attempts to answer the questions How did phosphorus atoms which are produced inside the inner cores of a handful of huge stars become concentrated in relatively high proportions in the organisms composing Earth's biosphere And how did these phosphate derivatives manage to be included in such a great variety of organic molecules playing essential biochemical roles in all known life forms Due to the interdisciplinary nature of the topic the volume is

arranged in three sections The first section introduces the fundamental concepts and notions of physics chemistry and biology necessary for the proper understanding of the topics discussed within an astronomical framework The author then focuses on the role of phosphorus and its compounds within the context of chemical evolution in galaxies considering its relevance in most essential biochemical functions as well as its peculiar chemistry under different physicochemical conditions The third section provides an overall perspective on the role of phosphorus and its compounds in current areas of research of solid state physics materials engineering nanotechnology or medicine

**Chemical Thermodynamics: Principles and Applications** J. Bevan Ott, Juliana Boerio-Goates, 2000-06-07 Chemical Thermodynamics Principles and Applications presents a thorough development of the principles of thermodynamics an old science to which the authors include the most modern applications along with those of importance in developing the science and those of historical interest The text is written in an informal but rigorous style including anecdotes about some of the great thermodynamicists with some of whom the authors have had a personal relationship and focuses on real systems in the discussion and figures in contrast to the generic examples that are often used in other textbooks The book provides a basic review of thermodynamic principles equations and applications of broad interest It covers the development of thermodynamics as one of the pre eminent examples of an exact science A discussion of the standard state that emphasizes its significance and usefulness is also included as well as a more rigorous and indepth treatment of thermodynamics and discussions of a wider variety of applications than are found in more broadly based physical chemistry undergraduate textbooks Combined with its companion book Chemical Thermodynamics Advanced Applications the practicing scientist will have a complete reference set detailing chemical thermodynamics Outlines the development of the principles of thermodynamics including the most modern applications along with those of importance in developing the science and those of historical interest Provides a basic review of thermodynamic principles equations and applications of broad interest Treats thermodynamics as one of the preeminent examples of an exact science Provides a more rigorous and indepth treatment of thermodynamics and discussion of a wider variety of applications than are found in more broadly based physical chemistry undergraduate textbooks Includes examples in the text and exercises and problems at the end of each chapter to assist the student in learning the subject Provides a complete set of references to all sources of data and to supplementary reading sources

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