

Chapter 10: Nuclear Reactions

Goals of Period 10

Section 10.1: To describe the half-life of nuclei and radio-carbon dating

Section 10.2: To further explore the relationship between mass and energy

Section 10.3: To calculate the binding energy of atomic nuclei

In Period 8 we discovered that unstable isotopes tend to decay into more stable nuclei. In Period 9 we found that unstable isotopes can decay by emitting alpha, beta, or gamma radiation, or by nuclear fission. Next we discuss the rate at which an isotope decays – the half life of the isotope. We then examine the effect on the mass and the energy of a nucleus when ionizing radiation is emitted.

10.1 Rates of Radioactive Decay

Half-Life

The decay of unstable, radioactive nuclei is a random process, and it is impossible to predict when a given nucleus will decay. Nuclei do not all decay at once, nor do those that have lived a long time show signs of aging like people do. The best one can do is to make the statistical statement that half the nuclei in a large sample of an unstable isotope will decay in a time characteristic of that isotope, called the **half-life** of the isotope and denoted by $T_{1/2}$. A similar statistical statement would be that in a large sample of coin tosses, half would come up heads. However, this does not mean that you get 5 heads in every 10 tosses nor does it tell you which of the coins will come up heads and which will come up tails.

In Period 1 we discussed exponential growth and decay. In exponential growth, a quantity doubles its size in a characteristic time called the doubling time. The equation describing exponential growth is

$$N = B \times 2^t \quad \text{(Equation 10.1)}$$

where

N = the amount of the quantity at a given time

t = the number of time periods elapsed

B = the initial amount of the quantity

In exponential decay, a quantity decreases by one half in a characteristic time, called the half life. A radioactive isotope is described by exponential decay, in which (statistically) the number of nuclei of the isotope is cut in half in a characteristic time, $T_{1/2}$. The equation for exponential decay is

$$N = B \times 2^{-t} = B \times \frac{1}{2^t} \quad \text{(Equation 10.2)}$$

Table 10.1 uses this relationship to compare the number of half lives elapsed to the portion of a sample remaining.

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JL Elias



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Nuclear Reactions W. M. Gibson, 1971 *Modern Physics* Paul Allen Tipler, Ralph Llewellyn, 2003 Tipler and Llewellyn's acclaimed text for the intermediate level course not the third semester of the introductory course guides students through the foundations and wide ranging applications of modern physics with the utmost clarity without sacrificing scientific integrity.

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Experimental Nuclear Physics: Physics of atomic nucleus Konstantin Nikiforovich Mukhin, 1987

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Table of Contents Chapter 10 Nuclear Reactions

1. Understanding the eBook Chapter 10 Nuclear Reactions
 - The Rise of Digital Reading Chapter 10 Nuclear Reactions
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 10 Nuclear Reactions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 10 Nuclear Reactions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 10 Nuclear Reactions
 - Personalized Recommendations
 - Chapter 10 Nuclear Reactions User Reviews and Ratings
 - Chapter 10 Nuclear Reactions and Bestseller Lists
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- Chapter 10 Nuclear Reactions Budget-Friendly Options
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 - Chapter 10 Nuclear Reactions Compatibility with Devices
 - Chapter 10 Nuclear Reactions Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Chapter 10 Nuclear Reactions
 - Highlighting and Note-Taking Chapter 10 Nuclear Reactions
 - Interactive Elements Chapter 10 Nuclear Reactions
- 8. Staying Engaged with Chapter 10 Nuclear Reactions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 10 Nuclear Reactions
- 9. Balancing eBooks and Physical Books Chapter 10 Nuclear Reactions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 10 Nuclear Reactions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 10 Nuclear Reactions
 - Setting Reading Goals Chapter 10 Nuclear Reactions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 10 Nuclear Reactions
 - Fact-Checking eBook Content of Chapter 10 Nuclear Reactions
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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