

1. **Disagreement** between the basic solute and solute. ClO_2 is made up of a chlorine and (two) weak oxygens. The solvent is purely the substance in question that is the chemical element. It is also in a substance that is dissolved in the solvent more or less. There is the solvent and ClO_2 is the solute in a self-solvent solution.
2. What is the concentration of solute in the solution of the following aqueous solution?
- a. $0.20 \text{ M Na}_2\text{SO}_4$ $\text{Na}_2\text{SO}_4 \rightarrow \frac{0.20 \text{ mol Na}_2\text{SO}_4}{1 \text{ L}} \times \frac{2 \text{ mol Na}^+}{1 \text{ mol Na}_2\text{SO}_4} = 0.40 \text{ M}$
- b. $0.20 \text{ M Na}_2\text{SO}_4$ $\text{Na}_2\text{SO}_4 \rightarrow \frac{0.20 \text{ mol Na}_2\text{SO}_4}{1 \text{ L}} \times \frac{2 \text{ mol SO}_4^{2-}}{1 \text{ mol Na}_2\text{SO}_4} = 0.40 \text{ M}$
- c. $0.20 \text{ M Na}_2\text{SO}_4$ $\text{Na}_2\text{SO}_4 \rightarrow \frac{0.20 \text{ mol Na}_2\text{SO}_4}{1 \text{ L}} \times \frac{2 \text{ mol Na}^+}{1 \text{ mol Na}_2\text{SO}_4} = 0.40 \text{ M}$
3. A solution is prepared by dissolving 25.0 g of sodium sulfate in enough water to prepare 250 mL of solution.
- a. What is the molarity of sodium sulfate in the solution?
- $$\text{M}_\text{Na}_2\text{SO}_4 = \frac{25.0 \text{ g Na}_2\text{SO}_4}{250.0 \text{ mL}} \times \frac{1 \text{ mol Na}_2\text{SO}_4}{142.04 \text{ g Na}_2\text{SO}_4} = 0.00176 \text{ mol/L}$$
- $$\text{M}_\text{Na}_2\text{SO}_4 = \frac{25.0 \text{ g Na}_2\text{SO}_4}{250.0 \text{ mL}} \times \frac{1 \text{ mol Na}_2\text{SO}_4}{142.04 \text{ g Na}_2\text{SO}_4} = 0.00176 \text{ mol/L}$$
- b. What are the molarities of the sodium and sulfate ions in the solution?
- $$\text{Na}^+ = \frac{0.00176 \text{ mol Na}_2\text{SO}_4}{1 \text{ L solution}} \times \frac{2 \text{ mol Na}^+}{1 \text{ mol Na}_2\text{SO}_4} = 0.00352 \text{ M}$$
- $$\text{SO}_4^{2-} = \frac{0.00176 \text{ mol Na}_2\text{SO}_4}{1 \text{ L solution}} \times \frac{1 \text{ mol SO}_4^{2-}}{1 \text{ mol Na}_2\text{SO}_4} = 0.00176 \text{ M}$$
- c. How many moles of sodium ions are present in 1 L of the solution?
- $$\text{Na}^+ = \frac{0.00352 \text{ mol Na}^+}{1 \text{ L}} \times 1.00 \text{ L} = 0.00352 \text{ mol Na}^+$$
4. A solution of Na_2SO_4 which has a molarity of 0.10 M and contains 0.0500 moles of potassium ions. What is the molarity of the Na_2SO_4 solution?
- $$\frac{0.0500 \text{ mol K}^+}{0.0500 \text{ L}} = \frac{0.0500 \text{ mol K}^+}{0.0500 \text{ L}} \times \frac{1 \text{ mol Na}_2\text{SO}_4}{2 \text{ mol K}^+} = 0.000500 \text{ M Na}_2\text{SO}_4$$
5. What are the freezing-point depression?
- Substance consists of cooling, mostly in the solvent and placing the solute particles in the solution.
6. Explain the boiling-point effect.
- One CH_3 group of negative compounds are composed and placed very poorly with water. Compounds with more CH_3 groups interact in poorly with water than they are called hydrophobic. When a hydrophobic compound is placed in water, the water molecules interact in a network that connects with the compound. As a result, hydrophobic compounds are not soluble in water.
7. Explain whether each of the following substances might be more soluble in H_2O or C_6H_6 .
- a. NaCl (ionic) b. $\text{C}_6\text{H}_5\text{OH}$ (polar) c. NaCl (ionic) (polar) d. C_6H_6 (nonpolar)
- Hydrogen bonding substances (H_2O and $\text{C}_6\text{H}_5\text{OH}$) and ionic substances (NaCl and C_6H_6) are more soluble in water than in a nonpolar solvent, while ionic substances (NaCl) are more soluble in ionic solvents.

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