

# Chapter 10: Chemical analysis

## Knowledge organiser

### Pure and impure

In chemistry, a **pure** substance contains a single element or compound that is not mixed with any other substances.

Pure substances melt and boil at specific temperatures.

An **impure** substance contains more than one type of element or compound in a **mixture**.

Impure substances melt and boil at a range of temperatures.

### Formulations

**Formulations** are examples of mixtures. They have many different components (substances) that make them up in very specific proportions (amounts compared to each other).

Scientists spend a lot of time trying to get the right components in the right proportions to make the most useful product.

Formulations include fuels, cleaning agents, paints, alloys, fertilisers, and foods.



### Testing gases

Common gases can be identified using the follow tests:

Gas	What you do	What you observe if gas is present
hydrogen	hold a lighted splint near the gas	hear a squeaky pop
oxygen	hold a glowing splint near the gas	splint re-lights
carbon dioxide	bubble the gas through limewater	the limewater turns milky (cloudy white)
chlorine	hold a piece of damp litmus near the gas	litmus turns the litmus white

### Chromatography

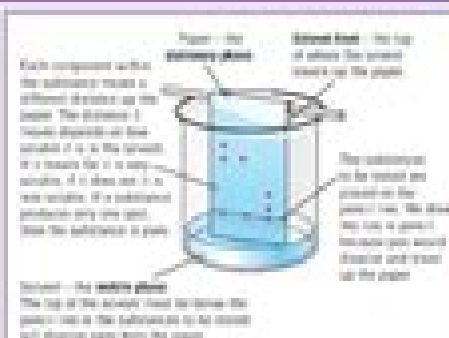
**Chromatography** is a method to separate different components in a mixture. It is set up as shown here, with a piece of paper in a beaker containing a small amount of solvent.

The **R<sub>f</sub> value** is a ratio of how far up the paper a certain spot moves compared to how far the **solvent** has travelled.

$$R_f = \frac{\text{distance moved for substance}}{\text{distance moved for solvent}}$$

It will always be a number between 0 and 1.

The **R<sub>f</sub> value** depends on the solvent and the temperature, and different substances will have different **R<sub>f</sub> values**. The **R<sub>f</sub> values** for particular solvents can be used to identify a substance.



### Separating mixtures

Mixtures can be separated by:

- Filtration** - separates insoluble solids from a liquid
- crystallisation** - evaporates a solvent (liquid) leaving the solute (solid)
- simple distillation** - separates solvent from a solution as long as the solvent has a lower boiling point than the solute
- fractional distillation** - separates two or more substances from a liquid mixture



#### Key terms

Make sure you can write a definition for these key terms.

chromatography

formulation  
solvent

impure  
solvent front

mobile phase  
stationary phase

pure

R<sub>f</sub> value

# Chapter 10 Study Guide Chemical Reactions Answers

**H. Stephen Stoker, Danny V. White**



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