

Chemistry

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(Chapter 1)(Solutions) (Exercises)
(Class 12)

Question 1.26:

If the density of some lake water is 1.25 g mL^{-1} and contains 92 g of Na^+ ions per kg of water, calculate the molality of Na^+ ions in the lake.

Answer 1.26:

Number of moles present in 92 g of Na^+ ions = $\frac{92 \text{ g}}{23 \text{ g mol}^{-1}} = 4 \text{ mol}$

Therefore, molality of Na^+ ions in the lake = $\frac{4 \text{ mol}}{1 \text{ kg}} = 4 \text{ m}$

Question 1.27:

If the solubility product of CuS is 6×10^{-16} , calculate the maximum molarity of CuS in aqueous solution.

Answer 1.27:

Solubility product of CuS , $K_{sp} = 6 \times 10^{-16}$ Let s be the solubility of CuS in mol L^{-1} .



$$K_{sp} = [\text{Cu}^{2+}][\text{S}^{2-}]$$

Now,

$$= s \times s = s^2$$

Then, we have, $K_{sp} = s^2 = 6 \times 10^{-16}$

$$\Rightarrow s = \sqrt{6 \times 10^{-16}} = 2.45 \times 10^{-8} \text{ mol L}^{-1}$$

Hence, the maximum molarity of CuS in an aqueous solution is $2.45 \times 10^{-8} \text{ mol L}^{-1}$.

Question 1.28:

Calculate the mass percentage of aspirin ($\text{C}_9\text{H}_8\text{O}_4$) in acetonitrile (CH_3CN) when 6.5 g of $\text{C}_9\text{H}_8\text{O}_4$ is dissolved in 450 g of CH_3CN .

Answer 1.28:

6.5 g of $\text{C}_9\text{H}_8\text{O}_4$ is dissolved in 450 g of CH_3CN .

Then, total mass of the solution = $(6.5 + 450) \text{ g} = 456.5 \text{ g}$

Therefore, mass percentage of $\text{C}_9\text{H}_8\text{O}_4 = \frac{6.5}{456.5} \times 100 \% = 1.424 \%$

Question 1.29:

Nalorphene ($\text{C}_{19}\text{H}_{21}\text{NO}_3$), similar to morphine, is used to combat withdrawal symptoms in narcotic users. Dose of nalorphene generally given is 1.5 mg. Calculate the mass of $1.5 \times 10^{-3} \text{ m}$ aqueous solution required for the above dose.

Answer 1.29:

The molar mass of Nalorphene ($\text{C}_{19}\text{H}_{21}\text{NO}_3$) is given as:

$$19 \times 12 + 21 \times 1 + 1 \times 14 + 3 \times 16 = 311 \text{ g mol}^{-1}$$

In $1.5 \times 10^{-3} \text{ m}$ aqueous solution of nalorphene, 1 kg (1000 g) of water contains $1.5 \times 10^{-3} \text{ mol}$

$$= 1.5 \times 10^{-3} \times 311 \text{ g} = 0.4665 \text{ g}$$

Therefore, total mass of the solution = $(1000 + 0.4665) \text{ g} = 1000.4665 \text{ g}$

This implies that the mass of the solution containing 0.4665 g of nalorphene is 1000.4665 g.

Therefore, mass of the solution containing 1.5 mg of nalorphene is: $\frac{1000.4665 \times 1.5 \times 10^{-3}}{0.4665} \text{ g} = 3.22 \text{ g}$

Hence, the mass of aqueous solution required is 3.22 g.

Chapter 1 Solutions

Joacim Rocklöv



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Table of Contents Chapter 1 Solutions

1. Understanding the eBook Chapter 1 Solutions
 - The Rise of Digital Reading Chapter 1 Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 1 Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 1 Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 1 Solutions
 - Personalized Recommendations

- Chapter 1 Solutions User Reviews and Ratings
- Chapter 1 Solutions and Bestseller Lists
- 5. Accessing Chapter 1 Solutions Free and Paid eBooks
 - Chapter 1 Solutions Public Domain eBooks
 - Chapter 1 Solutions eBook Subscription Services
 - Chapter 1 Solutions Budget-Friendly Options
- 6. Navigating Chapter 1 Solutions eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 1 Solutions Compatibility with Devices
 - Chapter 1 Solutions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 1 Solutions
 - Highlighting and Note-Taking Chapter 1 Solutions
 - Interactive Elements Chapter 1 Solutions
- 8. Staying Engaged with Chapter 1 Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 1 Solutions
- 9. Balancing eBooks and Physical Books Chapter 1 Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 1 Solutions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 1 Solutions
 - Setting Reading Goals Chapter 1 Solutions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 1 Solutions
 - Fact-Checking eBook Content of Chapter 1 Solutions

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web en 1907 se estrena Águila de blasón la primera de las tres comedias bárbaras escritas bajo el impacto que en valle inclán supuso la lectura de shakespeare el enfrentamiento de don juan manuel de montenegro consus hijos lleva al sórdido escenario de la galicia rural los ecos de el rey lear