

1. Suppose you have a question about some of the CO 202 course material that you want to ask your instructor. The best way to communicate with the instructor is by...

- a. Email
- b. Telephone
- c. Dropping by your instructors office
- d. All of the above.

Students who successfully complete this course will be able to...

- e. Distinguish positive and normative macroeconomic issues.
- f. Make appropriate choices among alternative frames of reference for analyzing macroeconomic issues.
- g. Use relevant information to analyze macroeconomic issues.
- h. All of the above

2. Students who successfully complete this course will be able to...

- a. Identify macroeconomic issues.
- b. Analyze the behavior of individual markets.
- c. Both of the above.
- d. None of the above.

3. The required material can be found in...

- a. The Macroeconomics in Context text.
- b. The Critical Thinking in Economics text.
- c. Both of the above.
- d. Mankiw's Principles of Macroeconomics text.
- e. All of the above are possible resources but none is required.

4. How will the quizzes and exams be taken?

- a. Online, at the course Canvas website.
- b. Through hardcopy exams.
- c. Either of the above.
- d. Online, at a third party's website.

5. What restrictions are there when taking quizzes and exams?

- a. Each quiz and each exam can only be taken once.
- b. Each quiz can be taken twice, but each exam can only be taken once.
- c. Each quiz can be taken as many times as desired before the due date.
- d. Each quiz and exam must be taken in a fixed amount of time.
- e. Each exam must be completed in 30 minutes.

6. A term paper is required for this course.

- a. True
- b. False

Eco 202 Fall Final Exam With Answers

Huangqi Zhang



Eco 202 Fall Final Exam With Answers:

Eco 202 Fall Final Exam With Answers Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the energy of words has been evident than ever. They have the ability to inspire, provoke, and ignite change. Such could be the essence of the book **Eco 202 Fall Final Exam With Answers**, a literary masterpiece that delves deep into the significance of words and their effect on our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book's key themes, examine its writing style, and analyze its overall impact on readers.

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web elements in the second row of the periodic table place their electrons in the 2n shell as well as the 1n shell for instance lithium li has three electrons two fill the 1 s orbital and the third is placed in the 2 s orbital giving an electron configuration of 1 s 2 2 s 1

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web one more point needs to be emphasized about the relationship between electron configuration and the periodic table the atoms of elements in the same vertical column of the table have similar electron configurations for example consider the alkaline earth elements group iia

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web the electron removed during the ionization of beryllium he 2s 2 is an s electron whereas the electron removed during the ionization of boron he 2s 2 2p 1 is a p electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are he 2 helium 1s 2 ne 10 neon 1s 2 2s 2 2p 6

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web q1 what is the electron configuration of the element potassium answer the atomic number of potassium is 19 a neutral potassium element has 19 electrons its electron configuration is $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$ or simply put as $4s^1$ q2 what is the total number of electrons that can be filled in s p and d subshells

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web by building up from hydrogen this table can be used to determine the electron configuration for any atom on the periodic table we will now construct the ground state electron configuration and orbital diagram for a selection of atoms in the first and second periods of the periodic table

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