Name : \_\_\_\_\_\_\_ Date: \_\_\_\_\_\_

# **Classifying Chemical Reactions**

- A. Classify the following reaction as synthesis, decomposition, single-replacement, double-replacement, and combustion:
- 1) Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> + 3 BaCl<sub>2</sub> --- 3 BaSO<sub>4</sub> + 2 AlCl<sub>3</sub> \_\_\_\_\_
- 2) Al<sub>2</sub>S<sub>3</sub> -- 2 Al + 2 S
- 3) 2 NaOH + CuSO<sub>4</sub> --- Na<sub>2</sub>SO<sub>4</sub> + Cu(OH)<sub>2</sub> \_\_\_\_\_
- 4) Fe + H<sub>2</sub>SO<sub>4</sub> → FeSO<sub>4</sub> + H<sub>2</sub> \_\_\_\_\_
- 5) C<sub>4</sub>H<sub>12</sub> + 7O<sub>2</sub> 4CO<sub>2</sub> + 6H<sub>2</sub>O \_\_\_\_\_
- B. Balance the following equations and indicate the type of reaction:
- 1)  $H_2S$  +  $O_2$   $\longrightarrow$   $SO_2$  +  $H_2O$  \_\_\_\_\_
- 2) C<sub>5</sub>H<sub>9</sub>O + O<sub>2</sub> → CO<sub>2</sub> + H<sub>2</sub>O \_\_\_\_\_
- 3) Al + NiBr₂ → AlBr₃ + Ni \_\_\_\_\_
- 4) AI + O₂ → AI₂O₃ \_\_\_\_\_
- 5) H<sub>2</sub>O<sub>2</sub> + O<sub>2</sub> \_\_\_\_\_\_

# **Cloze Ing In On Sciene Lesson 5 Chemical Reactions**

**M Woodhall** 

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