Outcomes

Upon completion of this lab, the student should be able to

- proficiently calculate molarities for solutions.
- prepare a solution of known concentration.
- prepare a dilute solution from a more concentrated one.
- perform serial dilutions.
- use volumetric and Mohr pipets and a volumetric flask.
- perform calculations using errors

Background

Solutions are an important part of chemistry. In this lab you will practice preparing solutions of different concentrations. The amount of solute that is dissolved in a given quantity of solvent is expressed as the concentration of the solution. A dilute solution contains only a small amount of solute in a given amount of solution. The units chemists use most often to describe concentration of solutions are molarity units. The molarity, M, (sometimes denoted as Concentration "C") of a solution is the number of moles of solute in one litre of solution. To determine the molarity of a solution, the following equation can be used:

Molarity (M) =
$$\frac{\text{moles of solute}}{\text{Litres of solution}}$$

Example 1: How would 500.0 ml. of a 0.6000 M NaCl solution be prepared?

Since NaCl is a solid, its mass must be measured on a balance. The first step is to calculate the mass of solute needed. The molar mass for NaCl is 58.44 g/mol.

Answer: Using a balance, measure out 17.53 g of NaCl into a small beaker and transfer it to a 500 -mL volumetric flask. A funnel may be necessary to transfer the solid into the flask. Rinse the beaker and funnel with a small amount of distilled water into the flask. Then add distilled water to the flask until it is about ½ full. Cap the flask and invert it several times to dissolve the solid. Next, add water until the liquid is just below the etched line on the neck of the flask. Bring the water to the line by adding the last few drops of water drop-by-drop using either a wash bottle or dropper. Again, cap and invert the flask several times to ensure proper mixing. The final volume of the solution is 500.0 mL. See Figure 1 on the next page.

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Table of Contents Cas Ch101 Lab Manual

- 1. Understanding the eBook Cas Ch101 Lab Manual
 - The Rise of Digital Reading Cas Ch101 Lab Manual
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Cas Ch101 Lab Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Cas Ch101 Lab Manual
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Cas Ch101 Lab Manual
 - Personalized Recommendations
 - Cas Ch101 Lab Manual User Reviews and Ratings
 - ∘ Cas Ch101 Lab Manual and Bestseller Lists
- 5. Accessing Cas Ch101 Lab Manual Free and Paid eBooks
 - o Cas Ch101 Lab Manual Public Domain eBooks

- Cas Ch101 Lab Manual eBook Subscription Services
- Cas Ch101 Lab Manual Budget-Friendly Options
- 6. Navigating Cas Ch101 Lab Manual eBook Formats
 - o ePub, PDF, MOBI, and More
 - Cas Ch101 Lab Manual Compatibility with Devices
 - Cas Ch101 Lab Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Cas Ch101 Lab Manual
 - Highlighting and Note-Taking Cas Ch101 Lab Manual
 - Interactive Elements Cas Ch101 Lab Manual
- 8. Staying Engaged with Cas Ch101 Lab Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Cas Ch101 Lab Manual
- 9. Balancing eBooks and Physical Books Cas Ch101 Lab Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Cas Ch101 Lab Manual
- 10. Overcoming Reading Challenges
 - o Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Cas Ch101 Lab Manual
 - Setting Reading Goals Cas Ch101 Lab Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Cas Ch101 Lab Manual
 - Fact-Checking eBook Content of Cas Ch101 Lab Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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